

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

Diphenyl[(phenylsulfonyl)methyl]- λ^5 -phosphanethione

Viktoria H. Gessner

Institut fuer Anorganische Chemie, Julius-Maximilians-Universitaet Würzburg, Am Hubland, 97074 Würzburg, Germany
Correspondence e-mail: vgessner@uni-wuerzburg.de

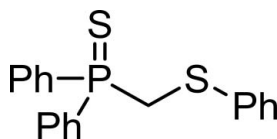
Received 21 February 2014; accepted 25 February 2014

Key indicators: single-crystal X-ray study; $T = 173$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.034; wR factor = 0.091; data-to-parameter ratio = 15.2.

The title compound, $\text{C}_{19}\text{H}_{17}\text{PS}_2$, results from the direct deprotonation of diphenylmethylphosphine sulfide and subsequent reaction with diphenyl disulfide. The C–P and C–S bond lengths of 1.8242 (18) and 1.8009 (18) Å, respectively, of the central P–C–S linkage are comparable to those found in the sulfonyl analogue, but are considerably longer than those reported for the dimetallated sulfonyl compound. The dihedral angle between the benzene rings of the diphenylmethyl moiety is $69.46(7)^\circ$. No distinct intermolecular interactions are present in the crystal structure.

Related literature

For the sulfonyl and dimetallated sulfonyl analogues, see: Schröter & Gessner (2012). For background to precursors for dilithio methandiides and their carbene complexes, see: Becker & Gessner (2014a,b); Cantat *et al.* (2006, 2008); Cavell *et al.* (2001); Cooper *et al.* (2010); Gessner (2012); Gessner *et al.* (2013); Harder (2011); Kasani *et al.* (1999); Liddle *et al.* (2011); Ong *et al.* (1999).



Experimental

Crystal data

 $\text{C}_{19}\text{H}_{17}\text{PS}_2$ $M_r = 340.42$

Monoclinic, $P2_1/c$
 $a = 9.3748(13)$ Å
 $b = 18.598(3)$ Å
 $c = 10.0941(14)$ Å
 $\beta = 101.044(2)^\circ$
 $V = 1727.3(4)$ Å³

$Z = 4$
Mo $K\alpha$ radiation
 $\mu = 0.39$ mm⁻¹
 $T = 173$ K
 $0.38 \times 0.18 \times 0.15$ mm

Data collection

Bruker APEX CCD diffractometer
Absorption correction: multi-scan
(SADABS; Bruker, 1999)
 $T_{\min} = 0.953$, $T_{\max} = 0.979$

10734 measured reflections
3033 independent reflections
2663 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.041$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.034$
 $wR(F^2) = 0.091$
 $S = 1.05$
3033 reflections

199 parameters
H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.31$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.23$ e Å⁻³

Data collection: SMART (Bruker, 2001); cell refinement: SAINT-Plus (Bruker, 1999); data reduction: SAINT-Plus; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 2012); software used to prepare material for publication: SHELXL97.

The author acknowledges the Deutsche Forschungsgemeinschaft, the Alexander von Humboldt Foundation and the Fonds der Chemischen Industrie for financial support.

Supporting information for this paper is available from the IUCr electronic archives (Reference: WM5008).

References

- Becker, J. & Gessner, V. H. (2014a). *Dalton Trans.* **43**, 4320–4325.
Becker, J. & Gessner, V. H. (2014b). *Organometallics*, doi:10.1021/om5001277.
Bruker (1999). *SADABS* and *SAINTE-Plus*. Bruker AXS Inc., Madison, Wisconsin, USA.
Bruker (2001). *SMART*. Bruker AXS Inc., Madison, Wisconsin, USA.
Cantat, T., Mézailles, P., Auffrant, A. & Le Floch, P. (2008). *Dalton Trans.* pp. 1957–1972.
Cantat, T., Ricard, L., Le Floch, P. & Mézailles, P. (2006). *Organometallics*, **25**, 4965–4976.
Cavell, R. G., Kamallesh Babu, R. P. & Aparna, K. (2001). *J. Organomet. Chem.* **617**, 158–169.
Cooper, O. J., Wooles, A. J., McMaster, J., Lewis, W., Blaker, A. J. & Liddle, S. T. (2010). *Angew. Chem. Int. Ed.* **49**, 5570–5573.
Farrugia, L. J. (2012). *J. Appl. Cryst.* **45**, 849–854.
Gessner, V. H. (2012). *Acta Cryst.* **E68**, o1045.
Gessner, V. H., Meyer, F., Uhrich, D. & Kaupp, M. (2013). *Chem. Eur. J.* **19**, 16729–16739.
Harder, S. (2011). *Coord. Chem. Rev.* **255**, 1252–1267.
Kasani, A., Babu, R. P. K., McDonald, R. & Cavell, R. G. (1999). *Angew. Chem. Int. Ed.* **38**, 1483–1484.
Liddle, S. T., Mills, D. P. & Wooles, A. J. (2011). *Chem. Soc. Rev.* **40**, 2164–2176.
Ong, C. M. & Stephan, D. W. (1999). *J. Am. Chem. Soc.* **121**, 2939–2940.
Schröter, P. & Gessner, V. H. (2012). *Chem. Eur. J.* **18**, 11223–11227.
Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.

supporting information

Acta Cryst. (2014). E70, o374 [doi:10.1107/S1600536814004292]

Diphenyl[(phenylsulfanyl)methyl]- λ^5 -phosphanethione

Viktoria H. Gessner

S1. Comment

Methylene compounds with two anion-stabilizing substituents, such as phosphonium or sulfonyl moieties, have found special interest as precursors for the corresponding dimetallated methandiides (Kasani *et al.*, 1999; Ong *et al.*, 1999; Cantat *et al.*, 2006; Cooper *et al.*, 2010). These dianions were found to be excellent ligands for the preparation of carbene complexes exhibiting a unique electronic structure (Gessner *et al.*, 2013). They allowed the synthesis of a variety of different complexes with early and late transition metals, but also with lanthanides and actinides (Cavell *et al.*, 2001; Cantat *et al.*, 2008; Harder, 2011; Liddle *et al.*, 2011).

As part of our studies on the synthesis of novel methandiides for the preparation of carbene complexes, we have developed a ligand and its dianionic analogue with a thiophosphoryl and sulfonyl sidearm (Schröter & Gessner, 2012). Thereby, the synthesis of a protonated precursor is best achieved *via* a two-step synthesis, with the first step being the lithiation of diphenylmethylphosphine sulfide and its reaction with diphenyl disulfide (Becker & Gessner, 2014*a,b*). This procedure furnishes the title compound, C₁₉H₁₇PS₂, in good yield. Oxidation of the sulfide to the sulfone gives way to the final ligand (Gessner, 2012).

The bond lengths and angles in the title compound are comparable to the sulfonyl analogue, but deviate considerably from the dimetallated compound (Schröter & Gessner, 2012). These differences are most pronounced in the P—C—S backbone. While the title compound features C—P and C—S distances of 1.8242 (18) and 1.8009 (18) Å, respectively, the sulfonyl substituted dianion shows C—P_{av} distances shortened by $\approx 7\%$ [1.710 (4) Å] and C—S distances shortened by $\approx 11\%$ [1.614 (3) Å] (Schröter & Gessner, 2012). Additionally, the P—C—S angle experiences a widening from 107.7 (1)° in the title compound to 121.4 (2)° in the methandiide. This is the result of a change in the hybridization of the central carbon atom from *sp*³ in the title compound to *sp*² in the methandiide.

No distinct intermolecular interactions (such as C—H \cdots S interactions) are present in the crystal structure of the title compound.

S2. Experimental

4.00 g (17.2 mmol) diphenylmethylphosphane sulfide were dissolved in 50 ml THF and 11.1 ml (17.2 mmol) *n*-butyllithium (1.55 M solution in hexane) were added drop-wise at 195 K. The mixture was allowed to warm to room temperature and stirred for additional 3 h. After cooling to 195 K 3.76 g (17.2 mmol) diphenyldisulfide dissolved in 30 ml THF were added. The mixture was stirred overnight, quenched by the addition of 50 ml water and extracted three times with 40 ml diethyl ether. After drying over sodium sulfate, the solvent was removed *in vacuo* and the residue purified by bulb-to-bulb distillation (oven temperature 473–483 K, 1x10⁻³ mbar). The title compound was obtained as a yellowish oil, which solidified after a couple of days (4.21 g, 12.3 mmol; 72%). ¹H NMR: (300.1 MHz, CDCl₃): δ = 3.92 (d, ²J_{PH} = 9.09 Hz, 2H; PCH₂S), 7.16–7.19 (m, 3H; CH_{SPh,meta/para}), 7.23–7.27 (m, 2H; CH_{SPh,ortho}), 7.40–7.47 (m, 6H; CH_{PPh,meta/para}), 7.72–7.79 (m, 4H; CH_{PPh,ortho}). ¹³C {¹H} NMR: (75.5 MHz, CDCl₃): δ = 38.5 (d, ¹J_{PC} = 52.25 Hz; PCH₂S), 127.1 (C_{SPh,ortho}),

128.6(d, $^3J_{PC} = 12.33$ Hz; $C_{\text{PPh},\text{meta}}$), 129.0 ($C_{\text{SPh},\text{meta}}$), 130.4 ($C_{\text{SPh},\text{para}}$), 131.0(d, $^1J_{PC} = 82.59$ Hz; $C_{\text{PPh},\text{ipso}}$), 131.6 (d, $^4J_{PC} = 10.24$ Hz; $C_{\text{PPh},\text{ortho}}$), 131.9 (d, $^3J_{PC} = 2.90$ Hz, $C_{\text{PPh},\text{para}}$), 135.5 (d, $^2J_{PC} = 5.78$ Hz; $C_{\text{SPh},\text{ipso}}$). $^{31}\text{P}\{^1\text{H}\}$ NMR: (122.0 MHz, CDCl_3): $\delta = 40.6$. Anal. Calcd for $\text{C}_{19}\text{H}_{17}\text{PS}$: C, 67.03; H, 5.03; S, 18.84; found: C, 66.80; H, 5.00; S, 19.09; GC—MS(ESI): tR = 17.18 min [353 K (2 min) -10 K min $^{-1}$ - 553 K (5 min)]; m/z (%): 340 (36) (M^+), 217 (100) $\{[\text{Ph}_2\text{PS}]^+\}$, 123 (46) $\{[\text{CH}_2\text{SPh}]^+\}$, 139 (81) $\{[\text{C}_6\text{H}_4\text{PS}]^+\}$.

S3. Refinement

The H atoms were refined on a riding model approximation in their ideal geometric positions with C—H = 0.95 Å for C—H(aromatic) and 0.99 for CH₂ atoms, respectively, with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

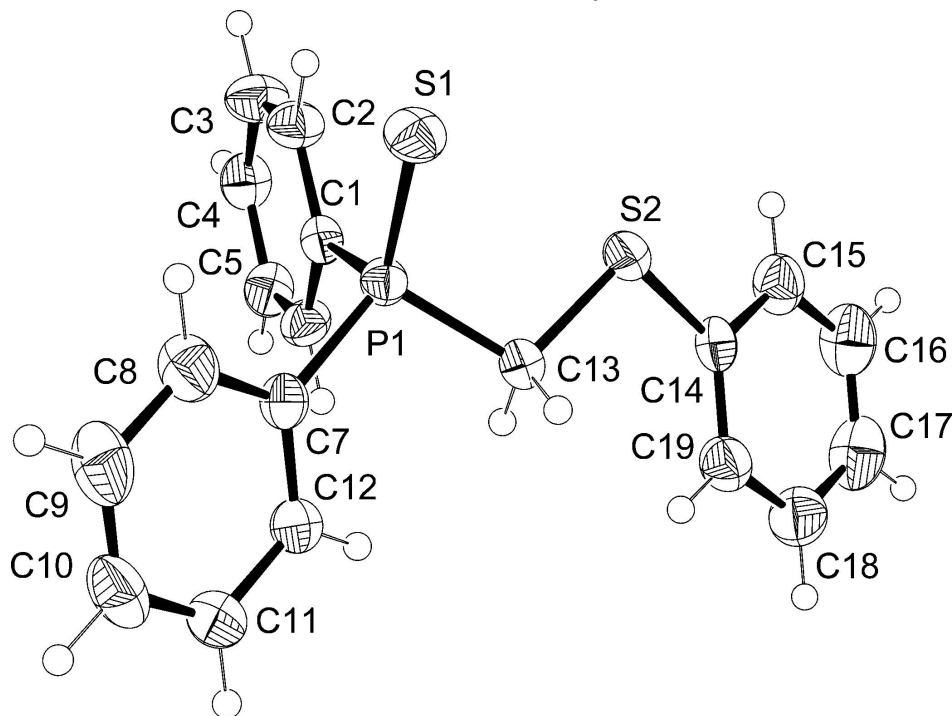


Figure 1

The molecular structure of the title compound. Displacement parameters are drawn at the 50% probability level.

Diphenyl[(phenylsulfanyl)methyl]- λ^3 -phosphanethione

Crystal data

$\text{C}_{19}\text{H}_{17}\text{PS}_2$
 $M_r = 340.42$
 Monoclinic, $P2_1/c$
 Hall symbol: -P 2ybc
 $a = 9.3748$ (13) Å
 $b = 18.598$ (3) Å
 $c = 10.0941$ (14) Å
 $\beta = 101.044$ (2)°
 $V = 1727.3$ (4) Å³
 $Z = 4$

$F(000) = 712$
 $D_x = 1.309$ Mg m $^{-3}$
 Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
 Cell parameters from 4352 reflections
 $\theta = 2.2$ – 25°
 $\mu = 0.39$ mm $^{-1}$
 $T = 173$ K
 Block, colourless
 $0.38 \times 0.18 \times 0.15$ mm

Data collection

Bruker APEX CCD diffractometer	10734 measured reflections
Radiation source: fine-focus sealed tube	3033 independent reflections
Graphite monochromator	2663 reflections with $I > 2\sigma(I)$
ω -scans	$R_{\text{int}} = 0.041$
Absorption correction: multi-scan (SADABS; Bruker, 1999)	$\theta_{\text{max}} = 25^\circ$, $\theta_{\text{min}} = 2.2^\circ$
$T_{\text{min}} = 0.953$, $T_{\text{max}} = 0.979$	$h = -11 \rightarrow 7$
	$k = -22 \rightarrow 22$
	$l = -12 \rightarrow 11$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.034$	H-atom parameters constrained
$wR(F^2) = 0.091$	$w = 1/[\sigma^2(F_o^2) + (0.0403P)^2 + 0.5374P]$
$S = 1.05$	where $P = (F_o^2 + 2F_c^2)/3$
3033 reflections	$(\Delta/\sigma)_{\text{max}} = 0.007$
199 parameters	$\Delta\rho_{\text{max}} = 0.31 \text{ e } \text{\AA}^{-3}$
0 restraints	$\Delta\rho_{\text{min}} = -0.23 \text{ e } \text{\AA}^{-3}$
0 constraints	
Primary atom site location: structure-invariant direct methods	

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
C1	0.52214 (19)	0.08028 (9)	0.30067 (17)	0.0288 (4)
C2	0.5712 (2)	0.02245 (10)	0.23596 (19)	0.0380 (4)
H2	0.535	0.0147	0.1425	0.046*
C3	0.6729 (2)	-0.02428 (11)	0.3069 (2)	0.0472 (5)
H3	0.7053	-0.0643	0.2623	0.057*
C4	0.7274 (2)	-0.01291 (10)	0.4421 (2)	0.0413 (5)
H4	0.7987	-0.0445	0.49	0.05*
C5	0.6782 (2)	0.04426 (10)	0.50795 (19)	0.0397 (5)
H5	0.7153	0.0518	0.6013	0.048*
C6	0.5750 (2)	0.09060 (10)	0.43818 (18)	0.0356 (4)
H6	0.5401	0.1295	0.484	0.043*
C7	0.2351 (2)	0.14772 (9)	0.28387 (18)	0.0311 (4)
C8	0.1148 (2)	0.10634 (12)	0.2316 (2)	0.0463 (5)
H8	0.1162	0.0763	0.1558	0.056*
C9	-0.0079 (2)	0.10857 (13)	0.2896 (2)	0.0551 (6)
H9	-0.0897	0.0796	0.2539	0.066*

C10	-0.0118 (2)	0.15236 (13)	0.3978 (2)	0.0510 (6)
H10	-0.0961	0.1538	0.4372	0.061*
C11	0.1064 (2)	0.19424 (11)	0.4494 (2)	0.0487 (5)
H11	0.1031	0.225	0.5239	0.058*
C12	0.2297 (2)	0.19198 (10)	0.3942 (2)	0.0402 (5)
H12	0.3114	0.2207	0.4314	0.048*
C13	0.4787 (2)	0.23097 (9)	0.24179 (18)	0.0322 (4)
H13A	0.4074	0.2696	0.2109	0.039*
H13B	0.5143	0.2366	0.3401	0.039*
C14	0.7380 (2)	0.30474 (9)	0.24404 (18)	0.0332 (4)
C15	0.8792 (2)	0.30812 (11)	0.2215 (2)	0.0467 (5)
H15	0.9118	0.2738	0.1642	0.056*
C16	0.9721 (2)	0.36083 (13)	0.2815 (3)	0.0568 (6)
H16	1.0679	0.3635	0.264	0.068*
C17	0.9273 (3)	0.40970 (13)	0.3668 (2)	0.0573 (6)
H17	0.9923	0.4458	0.4087	0.069*
C18	0.7882 (3)	0.40641 (12)	0.3914 (2)	0.0534 (6)
H18	0.7574	0.4402	0.4505	0.064*
C19	0.6925 (2)	0.35395 (11)	0.3302 (2)	0.0432 (5)
H19	0.5964	0.3518	0.3473	0.052*
P1	0.39369 (5)	0.14308 (2)	0.20578 (4)	0.02841 (14)
S1	0.34680 (6)	0.12132 (3)	0.01308 (5)	0.03941 (16)
S2	0.62814 (5)	0.23650 (3)	0.15364 (5)	0.03730 (15)

Atomic displacement parameters (Å²)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
C1	0.0287 (9)	0.0300 (9)	0.0276 (9)	-0.0035 (7)	0.0050 (7)	0.0025 (7)
C2	0.0430 (12)	0.0405 (10)	0.0295 (9)	0.0046 (9)	0.0047 (9)	-0.0024 (8)
C3	0.0533 (13)	0.0452 (11)	0.0428 (12)	0.0134 (10)	0.0090 (10)	-0.0047 (9)
C4	0.0389 (11)	0.0410 (11)	0.0422 (11)	0.0053 (9)	0.0031 (9)	0.0097 (9)
C5	0.0447 (12)	0.0409 (11)	0.0300 (10)	-0.0054 (9)	-0.0017 (9)	0.0033 (8)
C6	0.0440 (12)	0.0315 (9)	0.0303 (9)	-0.0023 (8)	0.0044 (9)	-0.0036 (7)
C7	0.0284 (10)	0.0311 (9)	0.0327 (9)	-0.0017 (7)	0.0035 (8)	0.0038 (7)
C8	0.0412 (12)	0.0577 (13)	0.0393 (11)	-0.0143 (10)	0.0055 (10)	-0.0050 (9)
C9	0.0346 (12)	0.0731 (16)	0.0558 (14)	-0.0193 (11)	0.0038 (11)	0.0046 (12)
C10	0.0346 (12)	0.0627 (14)	0.0595 (14)	0.0025 (10)	0.0187 (11)	0.0135 (11)
C11	0.0469 (13)	0.0447 (11)	0.0607 (13)	0.0000 (10)	0.0259 (11)	-0.0029 (10)
C12	0.0376 (11)	0.0376 (10)	0.0485 (11)	-0.0076 (9)	0.0158 (9)	-0.0062 (9)
C13	0.0333 (10)	0.0323 (9)	0.0326 (9)	-0.0035 (7)	0.0102 (8)	0.0012 (7)
C14	0.0336 (10)	0.0336 (9)	0.0322 (9)	-0.0029 (8)	0.0054 (8)	0.0091 (7)
C15	0.0374 (12)	0.0456 (12)	0.0589 (13)	-0.0007 (9)	0.0140 (10)	0.0068 (10)
C16	0.0345 (12)	0.0594 (14)	0.0737 (16)	-0.0078 (10)	0.0032 (12)	0.0103 (12)
C17	0.0520 (15)	0.0561 (14)	0.0552 (14)	-0.0188 (11)	-0.0111 (12)	0.0049 (11)
C18	0.0675 (16)	0.0501 (13)	0.0414 (12)	-0.0115 (11)	0.0075 (11)	-0.0078 (10)
C19	0.0445 (12)	0.0440 (11)	0.0430 (11)	-0.0064 (9)	0.0132 (10)	-0.0019 (9)
P1	0.0297 (3)	0.0295 (2)	0.0257 (2)	-0.00220 (18)	0.0045 (2)	-0.00026 (17)
S1	0.0459 (3)	0.0451 (3)	0.0251 (3)	0.0004 (2)	0.0013 (2)	-0.00118 (19)

S2	0.0409 (3)	0.0377 (3)	0.0369 (3)	-0.0078 (2)	0.0167 (2)	-0.00314 (19)
----	------------	------------	------------	-------------	------------	---------------

Geometric parameters (Å, °)

C1—C2	1.382 (3)	C11—C12	1.377 (3)
C1—C6	1.395 (2)	C11—H11	0.95
C1—P1	1.8138 (18)	C12—H12	0.95
C2—C3	1.384 (3)	C13—S2	1.8009 (18)
C2—H2	0.95	C13—P1	1.8242 (18)
C3—C4	1.379 (3)	C13—H13A	0.99
C3—H3	0.95	C13—H13B	0.99
C4—C5	1.379 (3)	C14—C19	1.385 (3)
C4—H4	0.95	C14—C15	1.387 (3)
C5—C6	1.384 (3)	C14—S2	1.7729 (19)
C5—H5	0.95	C15—C16	1.373 (3)
C6—H6	0.95	C15—H15	0.95
C7—C8	1.384 (3)	C16—C17	1.371 (3)
C7—C12	1.393 (3)	C16—H16	0.95
C7—P1	1.8137 (18)	C17—C18	1.375 (3)
C8—C9	1.388 (3)	C17—H17	0.95
C8—H8	0.95	C18—C19	1.387 (3)
C9—C10	1.369 (3)	C18—H18	0.95
C9—H9	0.95	C19—H19	0.95
C10—C11	1.373 (3)	P1—S1	1.9527 (7)
C10—H10	0.95		
C2—C1—C6	119.37 (17)	C11—C12—H12	119.9
C2—C1—P1	119.93 (14)	C7—C12—H12	119.9
C6—C1—P1	120.68 (13)	S2—C13—P1	107.72 (9)
C1—C2—C3	120.18 (17)	S2—C13—H13A	110.2
C1—C2—H2	119.9	P1—C13—H13A	110.2
C3—C2—H2	119.9	S2—C13—H13B	110.2
C4—C3—C2	120.23 (19)	P1—C13—H13B	110.2
C4—C3—H3	119.9	H13A—C13—H13B	108.5
C2—C3—H3	119.9	C19—C14—C15	119.44 (18)
C5—C4—C3	120.05 (18)	C19—C14—S2	125.31 (15)
C5—C4—H4	120	C15—C14—S2	115.23 (15)
C3—C4—H4	120	C16—C15—C14	120.3 (2)
C4—C5—C6	120.05 (17)	C16—C15—H15	119.9
C4—C5—H5	120	C14—C15—H15	119.9
C6—C5—H5	120	C17—C16—C15	120.4 (2)
C5—C6—C1	120.09 (17)	C17—C16—H16	119.8
C5—C6—H6	120	C15—C16—H16	119.8
C1—C6—H6	120	C16—C17—C18	119.9 (2)
C8—C7—C12	118.81 (18)	C16—C17—H17	120
C8—C7—P1	118.95 (15)	C18—C17—H17	120
C12—C7—P1	122.23 (14)	C17—C18—C19	120.4 (2)
C7—C8—C9	120.2 (2)	C17—C18—H18	119.8

C7—C8—H8	119.9	C19—C18—H18	119.8
C9—C8—H8	119.9	C14—C19—C18	119.6 (2)
C10—C9—C8	120.3 (2)	C14—C19—H19	120.2
C10—C9—H9	119.8	C18—C19—H19	120.2
C8—C9—H9	119.8	C7—P1—C1	108.53 (8)
C9—C10—C11	119.9 (2)	C7—P1—C13	103.53 (8)
C9—C10—H10	120.1	C1—P1—C13	104.55 (8)
C11—C10—H10	120.1	C7—P1—S1	113.32 (6)
C10—C11—C12	120.5 (2)	C1—P1—S1	113.12 (6)
C10—C11—H11	119.7	C13—P1—S1	112.99 (6)
C12—C11—H11	119.7	C14—S2—C13	102.51 (9)
C11—C12—C7	120.20 (19)		
